MASS TRANSFER FROM DROPLETS FALLING THROUGH A QUIESCENT BATH

BY
UDAY CHATTERJEE

DEPARTMENT OF METALLURGICAL ENGINEERING
INDIAN INSTITUTE OF TECHNOLOGY KANPUR
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UDAY CHATTERJEE

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CERTIFICATE

This is to certify that the study of "Masstransfer from droplets falling through a quiescent bath" has been carried out by Sri Uday Chatterjee under my supervisions and that this work has not been submitted elsewhere for a degree.

A Ghosh

(A. GHOSH)
ASSOCIATE PROFESSOR
DEPT. OF METALLURGICAL ENGG.
Indian Institute of Technology Kanpur

ABSTRACT

Transfer of indium from falling mercury-indium amalgam droplets to a static bath of aqueous ferric chloride solution was investigated at 35°C. The viscosity of the aqueous solution was varied by addition of glycerol upto 70 wt percent. The reaction was:

$$[In] + 3(Fe^{3+}) = 3(Fe^{2+}) + (In^{3+})$$

Two droplet sizes were employed. The change of concentration of In^{3+} in the aqueous phase as a function of time was followed by sampling and polarographic analysis. The diffusion coefficient of ferric ion at various concentrations were measured by polarographic technique. The terminal velocity and residence times of the droplets were measured by photographic technique. Efforts were also made to get an idea of the droplet oscillation during their fall through aqueous phase. Kinetic considerations and masstransfer calculations confirm aqueous control. However, log $\frac{(Fe^{3+})_t}{(Fe^{3+})_{t_0}}$ show some unexpected deviation from linearity when plotted against time. Variation of diffusion coefficient with concentration of ferric ion does not explain this departure from linearity completely. The drag coefficient VS Reynold's number behaviour is very close to that obtained by Calderbank et al and shows departure from solid sphere behaviour at high Reynolds number. The mass-transfer coefficient is one order magnitude higher than that expected from the solid sphere behaviour presumably due to droplet oscillation and possibly Marangani effect. The droplets attained terminal velocity at a very short distance after the fall.

ACKNOWLEDGEMENTS

It has been a great pleasure to work on the present problem concerning the most important but hither to over looked aspect of desulfurisation in Blast furnace. I am thankful to Dr. A. Ghosh and Dr. A.K. Biswas for eliciting this important and sensitive problem of today, and their endeavour and determination to meet the challenges encountered in successful completion of the work.

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1. INTRODUCTION

Since the development of two film mass transfer theory[1], Chemical Engineers have undertaken many fundamental studies on masstransfer from a drop phase to a surrounding liquid phase under different hydrodynamic conditions, various models have been developed which may be extended to specific metallurgical systems also.

In any Blast furnace operation, iron-ore is reduced to iron and while coming down through the stack it picks up Si, Mn, S, P etc. from the burden. The molten iron in the form of droplets of different sizes passes through the bosch region and collects in the hearth. The slag that forms from the gangue materials collects at the top of the metal. In normal operation the quantity of slag is about 50-70% of the total iron tapped. So, it is expected that the metal is always covered with a thick layer of slag and all the droplets of iron have to cross the thick viscous layer before they are collected in the hearth.

The existence and formation of the droplets inside the Blast furnace has been conclusively proved by many authors in the past. Seshadri[2] while explaining the "Mechanism of the formation of iron shots in B.F and low shaft furnace slag" had given a good review of the previous work done in this field. Kozakevitch[3] explained

the mechanism of desulphurization under R.F condition on the basis of emulsion and foam theory and proved that mass transfer from droplets occurs in the slag phase.

1.1 The mass transfer between a single drop and the surrounding field will depend on the

i) Shape and the nature of the drop

ii) Motion of the drop in the medium

iii) Diffusional flux on the surface.

1.1a Shape and nature of the drop

There are many situations when the liquid droplet passing through the continuous phase can not maintain its original shape. The shape will be ellipsoid and will oscillate along the major axis. Large drops can be considered to be oblate spheroid which may even oscillate to approximate prolate spheroidal form[4]. The oscillating drops show far greater rate of masstran - sfer than any other type. Saito[5] observed that when the drop density is very high and the ratio of the drop viscosity to the aquous phase viscosity is very small, then the drop may assume the prolate shape. The ecentri- city ratio E^1 as expressed by the author is

$$E^1 = \frac{\rho_D}{3}\left(1+\frac{\mu_D}{\mu_c}\right) - \frac{\rho_c}{3}\left\{\frac{10}{3}+\frac{319}{30}\left(\frac{\mu_D}{\mu_c}\right)+\frac{37}{5}\left(\frac{\mu_D}{\mu_c}\right)^2+\frac{1}{20}\left(\frac{\mu_D}{\mu_c}\right)^3\right\} \quad \text{---} (1)$$

When E^1 is greater than 0, the distortion is prolate and if $E^1 < 0$, the distortion is oblate.*

* Vide the list of nomenelatures given at the end of the text for explanation of the terms.

As the drop-size increases, the eccentricity of the nonoscillating ellipsoidal drop increases. Garner and Coworker[5] proposed the equation for calculating the area of such nonoscillating drop.

$$A = \frac{\pi}{2}\left[d_H^2 + \frac{d_V \cdot d_H}{E^2-1}\cdot \ln\left(E+\sqrt{E^2-1}\right)\right] \qquad (2)$$

The ratio of the area of an ellipsoid tothat of a sphere of equal volume is

$$\frac{A}{A_\Theta} = \frac{1}{2}\left[E^{2/3} + \frac{1}{E^{1/3}\sqrt{E^2-1}}\cdot \ln\left(E+\sqrt{E^2-1}\right)\right] \qquad (3)$$

If the drop of low viscosity moves through a high viscous field a series of shape change is expected. The shape of the droplet affects the Sherwood number in a complex way, but the masstransfer co-efficient remains proportional to the (diffusivity)$^{\frac{1}{2}}$ and agrees well with the theoretical analysis of transfer across mobile interfaces[4].

When the size of a drop falling in a low viscosity medium is increased beyond the laminar flow region, the drop begins to oscillate. The term oscillation refers to periodic changes from oblate ellipsoid to prolate and back to original form along the axis of symmetry. Banchero and Coworker[6] developed the equation by using the primary mode of oscillation and expressed as:

$$W \neq \omega = \left[\frac{192\,\sigma\,g_c}{(3\rho_D+2\rho_c)\,d_e^3}\right]^{1/2} \qquad (4)$$

1.1 b. Motion of the drop

The motion of the liquid drop in the liquid medium has been studied extensively by Chemical Engineers as a part of liq-liq. extraction technique. The motion of liquid mercury droplets in electrolytic solution is characterised by the existence of charge on the liquid surface. The motion of liquid droplets in the liquid medium is closely tied up with the size of the diffusional flux to the drop solution interface.

The effect of surface active material on the motion of drop had been studied by Levich and Coworkers.[7] It was observed that surface active materials may retard the motion on the surface of the drop when drop size is small. The surface active materials cannot retard the surface motion of a relatively large drop. Therefore, even in a medium containing such material, a large sized drop falls under the condition of unretarded surface motion. A model had been proposed by Huang and Kintner[8] which accounted for the reduced mass transfer to drop falling in medium containing surface active agents. The reduction in the interfacial area and changes in velocity and circulation pattern were the causes of reduced mass-transfer. The theoretical and experimental values agreed with each other well within 10%.

For a spherical liquid drop moving in another liquid field, the boundary does not remain rigid. It moves from front to rear stagnation along the axis of symmetry creating continuously a new area. Rubcznski and Hadamard[7] had shown that a change in boundary condition on the surface of a drop leads to a significant change in the velocity of the falling drops. For the fall of a small liquid drop in a liquid medium under the gravitational force, U, the terminal settling velocity had been expressed as

$$U = \frac{2}{3} \frac{g r^2 (\rho_D - \rho_c)}{\mu_c} \cdot \frac{\mu_c + \mu_D}{2\mu_c + 3\mu_D} \quad \text{cm/sec.} \qquad (5)$$

When the viscosity of the interior liquid is large, $\mu_D >> \mu_c$ the equation reduces to:

$$U = 2/9 \; g r^2 \frac{(\rho_D - \rho_c)}{\mu_c} \qquad (6)$$

which is Stoke's law. The Stokes law correction factor for liquid drop had been calculated from analysis of Hadamard and Rubcznski[5] and expressed as:

$$U_L = \left(\frac{3\mu_D + 3\mu_c}{3\mu_D + 2\mu_c} \right) U_s \qquad (7)$$

For a low viscosity liquid drop falling in high viscosity liquid field, the correction factor of 1.5 had been determined for a fully circulating drop. If the viscosities of the two fields are equal the correction factor of 1.2 may be considered.[5] Very often, inspite

of low viscosity of the liquid drop, the Stokes law is applied because of the rigidity of the surface of the drop due to the presence of unavoidable impurities.[9]

Velocity pattern in drops falling through glycerine solution had been studied by Garner and Haycock[10] and their observations revealed that no circulation was possible until the fall velocity exceeded 0.5 Cm/sec.

Due to mobility of the interface, the velocity gradients present in the liquid are smaller than those in the case of solid interface. With the same driving force (gravity), the velocity of steady fall of a liquid drop should be greater than the velocity of fall of a solid sphere.[7]

The liquid on the surface of the drop moves at a velocity

$$(U_\theta)_{r=a} = U_o \sin\theta \qquad (8)$$

Where U_o is the magnitude of liquid velocity at the drop equator and U_θ is the velocity on the surface of the drop at an angle θ from the drop equator. Again,

$$U_o = \frac{\mu_c}{2} \frac{U}{\mu_c + \mu_D} \qquad (9)$$

Where $\mu_D >> \mu_c$, the velocity on the surface is small and limit becomes zero in the case of solid surface.

1.1 C. Diffusional flux on the drop surface

Levich[7] deduced the diffusional flux to the surface of a drop, falling in a liquid where $R_e >> 1$. The density of the liquid in the drop was considered to be higher than that of the liquid medium. The total flux I, on the surface of the drop was found to be :

$$I = 8 (C_o - C) \sqrt{\frac{\pi}{5}} \left(\frac{D U_o}{r} \right)^{1/2} r^2 \qquad (10)$$

If $Re > 1$ and certain diffusion boundary layer exists on the surface of the droplet, its thickness δ can be found out from the relation :

$$U_o \frac{C}{r} \simeq D \frac{C}{\delta^2} \qquad (11)$$

For $Re > 1$, the velocity of the liquid U_o on the surface of the drop will be equal to the velocity of the falling drop. Hence by putting $U_o = U$ in equation(11) and by rearranging, the final form will be :

$$\delta = \sqrt{\frac{Dr}{U}} \qquad (12)$$

The general equation for diffusional flux on the surface of the drop is $I = \frac{D}{\delta} (C_o - C) A$ (13)
Substituting equation (12) in (13) and rearranging, the total flux on the drop surface in the case when $R_e > 1$ will be :

$$I = 4 \pi r^2 (C_o - C) \left(\frac{DU}{r} \right)^{1/2} \qquad (14)$$

The above equation is not valid for very high ~~Reynolds~~ Reynolds number.[7]

1.1 d. Masstransfer Models.

The masstransfer between a single sphere and surrounding fluid is usually divided into three parts.[11]

(a) Transfer within the sphere.

(b) Transfer through the interfacial phase

(c) Transfer outside the sphere.

The conditions which are necessary to be fulfilled prior to the studies are:

(a) Interfacial tension of the liquid drop should be high inorder to maintain the shape of the drop.

(b) The velocity of the continuous phase should be uniform in the region surrounding the drop.

(c) Fluids are dilute and binary diffusion equations are applicable.

(d) Heat generated by solution of transferring species are small.

(e) Pressure and temperature gradient is negligible.

When the drop behaves as a rigid sphere, the transfer rate can be determined from the molecular diffusivity, whereas some enhanced value other than molecular diffusivity has to be considered when the drops

are in the state of internal motion. The continuous phase masstransfer resistance follows closely that reported for solid sphere in the circulating regime where as the effective diffusivity within the drop is a constant multiple of the molecular diffusivity.[12]

It is possible in some cases that the masstransfer resistance is in the drop phase. The overall transfer rate will be controlled by the extraction mechanism inside the drop and to a [illegible] extent by the hydrodynamics of the system. Both molecular and convective diffusivity may influence the transfer rate in the dispersed phase. Extraction efficiency can be expressed in terms of Newman's equation,[13]

$$Ed = 1 - \frac{6}{\pi^2} \sum_{n=1}^{\infty} \frac{1}{n^2} \exp\left(\frac{-4\pi^2 n^2 D_D t}{d^2}\right) \quad (14)$$

Ed = fraction extracted (dimensionless).

The circulation model of Kronig and Brink[14] is valid when the relative motion of the drops induces circulation. Turbulence of high intensity in the continuous phase may cause turbulent movement inside the drop. Handlos and Baron[15] proposed the model for masstransfer under the above mentioned situation. The results from the series of experiment indicated that they do not agree well with the above models. So one experimental factor R has been introduced in the equation. R when multiplied

with molecular diffusivity becomes more or less the effective diffusivity.[13]

Garner and Coworker[16] studied the effect that droplet oscillation had on the continuous phase resistence. Rose and Kintner[16] suggested a model for vigorously oscillating single liquid drop, moving in a liquid field with the concept of interfacial stretch and internal droplet mixing. To predict the initial zone thickness for a spherical drop with uniform internal concentration of solute, two film theory has been used. The liquid film thickness is predicted by using the empirical correlation of Garner and Coworker which is:

$$N_{Sh_C} = 0.6\, N_{R_e}^{\frac{1}{2}}\, N_{S_c}^{\frac{1}{3}} \qquad (15)$$

To predict the inside film thickness, the penetration theory is used with contact time equal to the time for one oscillation cycle. This contact time was chosen as a result of photographic study of the pattern of movement inside a falling oscillating drop. During each time, the droplet underwent a period of oscillation, the interior of the drop was violently mixed, so the interface will be renewed during each droplet oscillation. Hence,

$$K_D = .45\, (D_D W)^{\frac{1}{2}} \text{ Cm/sec.} \qquad (16)$$

W = frequency of oscillation in radians/sec.

Hence the over all masstransfer coefficient is

$$\frac{1}{K_D} = \frac{1}{k_D} + \frac{m}{k_c} \quad \text{Where } m = \frac{(\rho_A - \rho_A^*) \text{ drop phase}}{(\rho_{A_0} - \rho_A) \text{ cont phase}} \qquad (17)$$

In Rose and Kintner model, one oscillation has been considered as surface life time and the area change resulting from one oscillation has been assumed constant. Angelo and Lightfoot[17] modified the Rose and Kintner concept and studied the penetration theory for surface stretch applied to oscillating droplets in the Reynolds number ranges between 200-2000. Further modification and studies in the Reynolds number range between 10-20 was conducted by Sy and Lightfoot[18] and they observed that the distortion of drop changes the Nusselts number by about 20%. All these studies have been made with liquids of low density and viscosity.

It has been known for quite sometime that the composition of the liquid metal at the bosch is very much different as compared to the metal at the hearth. Earlier workers tacitly assumed that the site for slag-metal reaction was the interface between the slag and the pool of metal. Therefore, they essentially simulated the slag and metal pool in the laboratory and studied the reaction Kinetics. However, from masstransfer considerations, there was no reason to ignore the chemical reaction while the metal droplets are passing through the slag. As a matter of fact, some recent studies[19][20] have shown

that the desulphurisation takes place primarily while droplets fall through the slag layer. When this work was undertaken no investigation in this direction could be located in the literature.

The situation is very complex involving thousands of droplets of different sizes and specific chemical and physical conditions. A clear understanding of the process may emerge through controlled drop studies. The knowledge gained from the performances of individual large drops may not be immediately applied to the complex operation. However, this is expected to be of considerable help to understand the masstransfer situation. It is in this context, the decision was taken to examine the phenomena exhibited by single drop moving through the viscous liquid layer with the help of a cold model.

2. PLAN AND OUTLINE OF THE WORK

The difficulties of handling molten slag and metal at high temperature can be eliminated with the help of room temperature systems in which the densities and viscosities of the two phases should bear some resemblance to the actual process. Different techniques had been adopted to study the masstransfer to or from the droplets and the bubbles. Constant volume experimental technique as adopted by Calderbank[22][21] and coworker to study the mass-transfer from bubbles can be applied to droplets also. Scintillation technique for instantaneous measurement of exchange rate of the solute on freely suspended single drop had helped in studying the exchange process without affecting the hydrodynamics of the system[23]. Calderbank[24] and Coworkers observed that mercury drops formed one drop diameter above the continuous phase were not distorted while passing through the liquid surface.

In this study, mercury drops were formed above the surface of the aqueous phase in conformity with the actual Blast furnace condition and the (1) Mass transfer from the droplets while passing through the aqueous phase (2) Diffusion co-efficient of the solute in the aqueous phase (3) Velocity and the residence time of the droplets in the aqueous phase were determined. The dropping rate was

so adjusted that only one drop at a time passed through the aqueous phase.

2.1 Determination of Masstransfer Co-efficient.

In the experiments, mercury droplets were used to simulate the metal phase and water with its viscosity adjusted with glycerol was equivalent to the slag phase. Typical reactions which are conveniently used for studying mass transfer between slag and metal without production of gas at the interface are:

$$2\,[In] + 3\,(Hg_2^{2+}) = 2(In^{3+}) + 6\,[Hg] \qquad (18)$$

$$[In] + 3(Fe^{3+}) = (In^{3+}) + 3(Fe^{2+}) \qquad (19)$$

Where square brackets indicate the metal phase and round brackets the aqueous phase.

Subramanium and Richardson[25] studied the mass transfer across interface agitated by bubbles in both systems represented by equations (18)(19) under aqueous and metal controlled conditions.

For the process of masstransfer from the droplets, reaction (19) can be favourably used. In-Hg amalgam of different droplet sizes were passed through aqueous phase of different ferric ion and glycerol concentrations. In the series of experiments it was planned to verify whether aqueous control condition existed in the system

or not.

The small amalgam pool formed at the bottom of the apparatus was kept separated from aqueous solution in order to avoid any reaction between the two liquids.

2.2 Diffusion Co-efficient of Fe^{3+} ion.

During the course of experiment, the concentration of ferric ion in the aqueous phase is supposed to be reduced, hence the diffusion coefficient of ferric ion is expected to vary. So, it is essential to determine the diffusion co-efficient of ferric ion as a function of concentrations in the aqueous phase. Diffusion co-efficient of ferric ion of various concentrations under different glycerol content similar to experimental conditions for masstransfer study were carried out by Polarographic technique.

2.3 Velocity and residence time.

Knowledge of velocity and residence time for the droplets in the aqueous phase is important in order to determine the hydrodynamics and the masstransfer results. Photographic techniques were adopted for the measurements of terminal velocity and the residence time of the droplets in the aqueous phase.

3. REAGENTS

The aqueous solutions used for studying the reaction in equation (19) consisted of $FeCl_3$ at various concentrations ranging from 10-75 m.moles/liter and A.R. grade glycerol between 35-70 wt percent for adjusting the viscosities of the aqueous phase. The pH of all the solutions were adjusted to 1.4 by adding A.R grade HCl before every experiment. All aqueous solutions were deacrated prior to use by passing purified nitrogen for 2 hours.

The amalgams were prepared from triple distilled mercury and indium of 99.999 percent purity.

3.1 Analytical Technique.

Polarographic techniques were used for measuring the concentrations of indic ion and the diffusion co-efficients of ferric ions in the aqueous phase. The apparatus consists of a mercury reservoir fitted with a capillary tube at the bottom and a calomel electrode in one branch of U tube and the other branch is the sample container separated by porous disc. Mercury drops are the dropping electrode formed from the capillary of 0.03- .05 mm bore under constant pressure and time interval. Calomel electrode has been provided to eliminate the possibility of unknown or nonreproducible anode potential, hence it acts as reference electrode. In polarography actually

the current vs. E.M.F plot is recorded (26)

A concentration gradient is established between the dropping electrode surface and the bulk solution. At equilibrium, the rate of discharge of ions by current is equal to the rate of diffusion to the electrode surface. For relatively small values of current, the concentration over potential is small and as the applied E.M.F is increased beyond the decomposition potential, the current exhibits the normal rise, and a stage is reached at which the concentration overpotential increases rapidly and the current reaches a limiting value, hence this is the maximum rate at which particular ion can be discharged under the given experimental condition. As the concentration of the bulk solution increases, the limiting current also increases. The galvanometer is unable to follow the periodic growth and fall of the current of each individual drop, so the saw toothed waves are observed. Thus in measuring the diffusion current, the average of the galvanometer oscillation is to be considered (27).

In order to obtain true diffusion current of a substance, a correction must be made for the residual current. The most reliable method for making this correction is to evaluate in a separate polarogram the residual current of the supporting electrolyte alone. The value of the residual current at the particular half wave potential

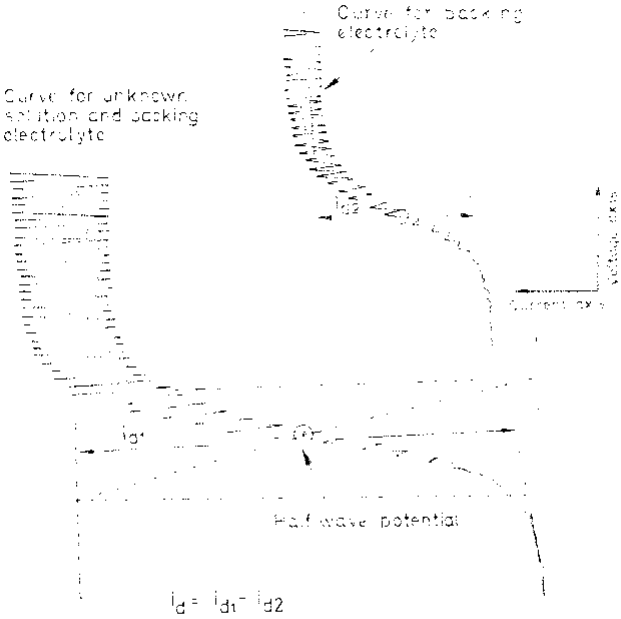


FIG. 1 CORRECTION TECHNIQUE FOR THE DETERMINATION OF CONCENTRATION AND DIFFUSION COEFFICIENT FROM POLAROGRAM

level of the corresponding element is subtracted from the total observed current[27]. Fig. 1 indicates the correction techniques for the determination of diffusion current. Correction technique has been adopted for the determination of concentration and diffusion co-efficients.

All the electroactive materials have their half wave potential. This is the potential at the point of current-voltage curve, half the distance between the residual current and the final limiting current. Hence it is customary to measure diffusion current at the half wave potential level. In the case of In, the value of half wave potential changes with pH. At pH 3.5- 4, half wave potential is -.62 volts[28] and -0.24 volts for Ferric ion at pH=4[29]. For the determination of concentration of In in the aqueous phase, the pH of the solution has to be maintained at 1.4 and with backing electrolyte, the resultant pH was within 3.5-4.

Before each experiment, the solution is deaerated by passing purified nitrogen for 10-15 minutes.

4. Description of the apparatus.

The apparatus used for masstransfer study is shown in Fig. 2. It is a 15 cm long tube of 4 cm diameter resting on a container for storing the used amalgam. Suitable siphoning arrangement has been provided to drain out the amalgam from the column at proper interval.

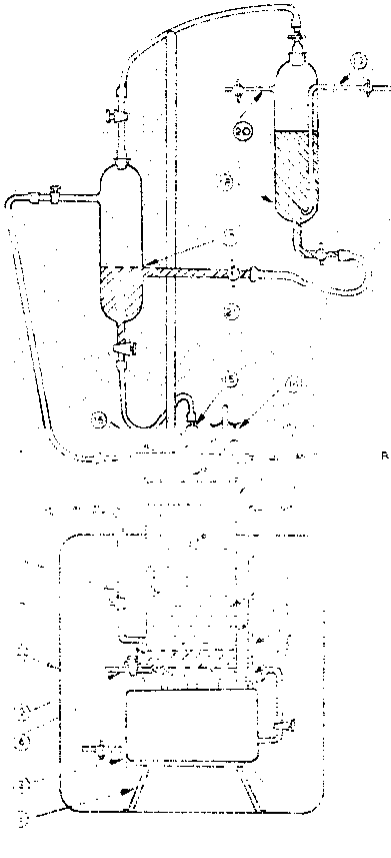


FIG.2 SCHEMATIC DIAGRAMME OF THE APPARATUS

Since the Hg-In amalgam gets oxidized very easily in air, suitable arrangement has been made to maintain nitrogen atmosphere on the liquid column. Since the deacration of the solution is required before each experiments, suitable provision has been made for this purpose. There are three openings at the top of the apparatus. The centre hole for dropping amalgam and the other two are for sampling and thermometer point. The whole apparatus is completely air tight and immersed in constant temperature bath.

The auxiliary part of the apparatus is the amalgam preparation and dropping unit. The amalgam is stored on separate column and drawn into the dropping column after certain interval of time in order to maintain more or less constant level in it. The height of the amalgam column from the top level of the amalgam pool to the tip of the capillary was maintained at 30 cm. Provisions are there to maintain nitrogen atmosphere on the amalgam storage and in the dropping unit.

The whole apparatus is made up by Pyrex glass and suitable ground glass joints has been used for sealing purposes.

4.1 Masstransfer Experiments.

Before starting the experiment, a layer of "Tetra Bromoethane" of 5 mm thickness was made at the bottom of the apparatus in order to separate the amalgam

pool from the aqueous solution. Tetra Bromo ethane is of density 2.6, reacts neither with aqueous phase nor with the amalgam . Hence, the masstransfer only from the droplets can be studied.

In all the experiments 200 cc of aqueous solution containing different proportions of glycerol and $FeCl_3$, after adjusting the pH to 1.4 was poured slowly on the top of the tetra Bromo ethane. The height of the aqueous column was 10 cm in all the cases. The solution was thoroughly deaerated by passing purified nitrogen for two hours.

Some experiments were conducted at 21-22°C in the present apparatus, where appreciable amount of transfer could not be detected. With a column height of 60 cm, sufficient amount of transfer was recorded with 60 cc aqueous solution at a temperature of 21-22°C. Since it was difficult to fulfill the experimental requirements in the tall column, experiments at higher temperature was tried. Hence all the experiments were conducted at 35°C. It was possible to control the temperature of the constant temperature bath upto $35 \pm .1$°C.

Capillaries of 1 mm and 0.2 mm inside diameter were used for dropping amalgam. All the capillary tips were polished and cleaned thoroughly before use.

Amalgam was prepared by pouring the weighed amount of mercury on the indium under nitrogen atmosphere. Blowing by purified nitrogen into the stored amalgam for 1-2 minutes was preferred to ensure thorough mixing. No scum formation on the surface of amalgam storage was detected. Storage column and the dropping unit remained under nitrogen atmosphere through out the experiment.

As soon as the preparation of amalgam was over, certain amount of amalgam was drawn out from storage to the dropping unit, deaerated the capillary and the connecting rubber hose. This amalgam was stored separately and subtracted from the total amount in order to calculate the actual amount of amalgam used in each experiment. Before starting the experiments, 8-10 drops of amalgam were collected in the weighing bottles under identical condition in order to determine the equivalent diameter of the droplets. Identical dropping rate could not be maintained in all the experiments but in individual experiments same drop rate from the beginning to the end was maintained. Drop rate varied between 125-155 per minute in the case of small capillary and 105-145 per minute for the bigger one. The above mentioned dropping rate was sufficient to ensure the passage of a single droplet through the liquid column at a time. The amount of amalgam passed through the solution varied between 1040-1370 gms in the case of small

capillary and 1830 - 2020 gm for the big one. The experiments with small capillary continued from 270-295 minutes and with big capillary from 82 - 110 minutes.

Subramanium and Richardsons[25] work indicated that the initial concentration of In in the amalgam had no effect on the observed rates of masstransfer. Two experiments were tried with different In concentrations and the same results were obtained. Hence it was decided to conduct all the experiments with In concentration of 0.06%.

Viscosity of the liquid phase varied from 0.73 cp to 11.75 cp and Ferric ion concentration from 10-75 m. moles per liter.

In all the experiments, after passing amalgam through the aqueous solution for certain interval of time, 2 cc of the solution was drawn out through the sampling hole and analysed polarographically for indium content. The characteristic half wave potential of In^{3+} and Fe^{2+} are widely separated. To prevent Fe^{3+} from oxidizing the dropping mercury electrode, 2M Potassium chloride containing 50 gm/liter of hydroxy amonium chloride was used as a backing electrolyte.[25] The addition reduced the ferric ion to Fe^{2+}. The diffusion current varied linearly with concentration and was not affected by glycerol content. Correction method as indicated earlier was followed to determine the net diffusion current for indic ion. Again

polarographically determined the diffusion current for indic ion of a standard indium chloride solution of known concentration in the presence of a backing electrolyte of 1 MKCl/liter. Similarly correction method was adopted in finding out the net diffusion current. Then from the relation

$$\frac{C_1}{C_2} = \frac{id_1}{id_2}$$

where C_1 = Concentration of indic ion in unknown solution in m. moles/liter. ~~(with moles/liter)~~

C_2 = Concentration of indic ion in standard solution of known concentration, in m.moles/liter.

id_1 = diffusion current for unknown solution in microampere.

id_2 = diffusion current for known solution in micro amp.

the concentration of indic ion in the aqueous phase was determined. This procedure was adopted in all the samples. By stoichiometry, from the concentration of indic ion, the amount of ferric ion consumed after certain interval of time was calculated.

4.2 Diffusion co-efficient of ferric Ion.

In the masstransfer experiments it was observed that the concentration of ferric ion in the aqueous phase was decreasing, hence the diffusion co-efficient values might also change. D_{Fe}^{3+} values are available only at 20°C. Since all the masstransfer experiments were carried out

at 35°C, so the D_{Fe}^{3+} values at different concentrations of ferric ion and glycerol were determined by polarographic technique at 35°C.

In the series of experiments, concentrations of ferric ion were varied between 6.25- 31.5 m. moles/liter in water, 12.5 - 31.5 in water + 35% wt glycerol and 12.5 - 64 m. moles per liter in water + 70 wt per cent glycerol media. The backing electrolyte[29] was 0.2M per liter sodium oxalate and .005% gelatine as maxima suppresser. It was observed that .005% gelatine was sufficient for water and water + 35 wt % glycerol medium but in case of 70 wt % solution, the gelatine addition was increased to .01 % . The pH of the solution was adjusted to 4 by adding acetic acid before each experiment.

Known volume of solution was taken in the sample holder and deaerated for 15 minutes before the start of the experiment. Temperature control was 35°C ± .5°C. The drop time and weight of mercury passed per second were determined. Correction method was adopted in determining the net diffusion current for each sample from the Polarograms. From the law governing the condition of diffusion at a dropping mercury electrode and periodic growth and fall of the droplets, theoritical equation for diffusion current as developed by Ilkovic[26][27] is:

$$i_d = 607\ F\ D_{Fe^{3+}}^{\frac{1}{2}}\ C\ m^{\frac{2}{3}}\ t^{1/6} \qquad (20)$$

id = diffusion current in micro ampere during the life of the drop.

F = number of Faraday electricity required per mole of the electrode reaction.

c = concentration in millimoles per liter.

m = rate of flow of mercury in m.g/sec.

t = drop time in sec.

$D_{Fe^{3+}}$ = diffusion co-efficient of ferric ion in Cm^2/sec. From eqn. (20).

$$D^{\frac{1}{2}}_{Fe^{3+}} = \frac{id}{607\, F.C\, m^{\frac{2}{3}} t^{\frac{1}{6}}} \qquad (21)$$

$D_{Fe^{3+}}$ for all the solution are calculated and reported in table no.3. The plot $D_{Fe^{3+}}$ vs concentration is shown in fig. 4 which helped in determining diffusion co-efficient at any concentration within the experimental range. The reproducibility of diffusion current values were checked and these were within ± 1 %.

4.3 Velocity and Residence time Determination.

The terminal velocity and the residence time of the droplets of different sizes in the different aqueous media similar to masstransfer experimental conditions were determined by photographic means. Several photographic exposures at regular time intervals of a drop falling through the aqueous phase were taken by using STROBOTAC- Type 1531-AB (General Radio Company) in a darkened room. For

bigger drops 8000 cycles/min and for the smaller one 4000 cycles/min r.p.m. settings were used. The photographs are shown in Fig. 3a,b. The photographic studies revealed that the small drops were reaching their terminal velocities within 1-1.5 cm travel distance from the liquid surface, the corresponding travel distance was 3-3.5 cm in the case of big drops. Residence time and the terminal velocities were calculated from drop intervals and the stroboscope cycle r.p.m values for all the cases and reported in table 1.b

Movie Camera was used in determining the change of shapes of the droplets which were taking place during its passage through the aqueous phase by using Stroboscope. It was possible to observe the changes in shape at different positions of the liquid column but due to certain limitations, complete shape changes could not be recorded hence shapes could not be measured. Some interesting movie photographs are shown in Fig. 4.

SMALL SIZE DROPLETS

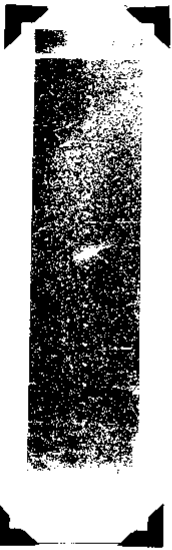
70 WT [illegible] GLYCEROL
+ $FeCl_3$

35 WT PERCENT GLYCEROL
+ $FeCl_3$

WATER + $FeCl_3$

Fig. 5a,b - Terminal velocity and residence time determination by Photographic technique.

BIG SIZE DROPLETS



35 WT PERCENT CLYCEROL
+ $FeCl_3$

WATER + $FeCl_3$

Fig. 3a,b- Terminal Velocity and residence time determination by Photographic technique.

Fig. 4.- Shows the oscillation and changes in drop size during its passage through the aqueous phase (water + PeHg).

5. [illegible]

Properties of the aqueous phase* and other experimental data are given in table 1a and 1b.

Table 1a

Exp. No.	Weight percent glycerol	Conc. of ferric ion in m. moles per liter	Viscosity in C.P	Density in gm/cc
1	70	25	11.75	1.175
2	70	75	11.75	1.19
3	70	10	11.75	1.17
4	70	25	11.75	1.175
5	water	25	.728	1
6	35	25	2.03	1.09
7	35	25	2.03	1.09
8	water	25	.728	1
9	water	25	.728	1

* Viscosity and density data taken from Handbook of Chemistry and Physics, Forty third edition.

TABLE 1b

Exp. No.	Drop dia in cm	Drop rate per min	Total weight of amalgam consumed in gms.	Terminal velocity in cm/sec	Resistance time in sec.
1	.159	125	1082.1	47.62	.225
2	.160	150	1370	47.62	.225
3	.159	135	1195.5	47.62	.225
4	.155	155	1220.7	47.62	.225
5	.157	155	1234	70.2	.167
6	.159	145	1172	52.5	.186
7	.284	145	2020.5	72.07	.163
8	.281	105	1834.5	73.50	.157
9	.158	135	1042	70.2	.167

Results from Mass transfer experiments are shown in table II

TABLE II

EXP.NO.	Conc. of in^{3+} in aqueous phase in m. moles per liter	Initial ferric ion conc.in m.moles per liter $(Fe^{3+})_{t_o}$	$\frac{(Fe^{3+})_t}{(Fe^{3+})_{t_o}}$	Time in sec.
	.14	25	0.9832	3000
	.276	25	0.9669	6000
1	.34	25	0.9592	9000
	.824	25	0.9010	16800
	.89	25	0.8933	17700
	1.0496	75	0.958	9000
2	1.4158	75	0.943	10500
	1.319	75	0.927	15000
	2.1787	75	0.912	17700
	.1263	10	0.962	6000
3	.3368 .5320	10 10	0.9390.8404	10500 15000
	.6481	10	0.8068	17700
	.1680	25	0.980	4500
4	.4620	25	0.944	9000
	.7638	25	0.908	13500
	1.1197	25	0.866	17100
	1.4033	25	0.8316	4500
5	2.989	25	0.6413	9000
	4.1047	25	0.5074	12600
	4.9117	25	0.4106	16200
	.8089	25	0.903	4500
6	1.733	25	0.792	9000
	2.629	25	0.6345	12600
	3.333	25	0.6000	16200
	.7086	25	0.915	1800
7	1.8099	25	0.783	3600
	2.5465	25	0.694	4500
	2.9255	25	0.649	4950
	.947	25	0.8864	1200
8	2.0625	25	0.7525	2700
	3.2622	25	0.6086	4500
	4.2160	25	0.4940	6600
	1.340	25	0.8392	4500
9	1.943	25	0.7668	9000
	3.5917	25	0.669	12600
	4.4473	25	0.4663	16200

Plots of $\frac{(Fe^{3+})_t}{(Fe^{3+})_{t_o}}$ vs time are shown in fig. 5a, 5b

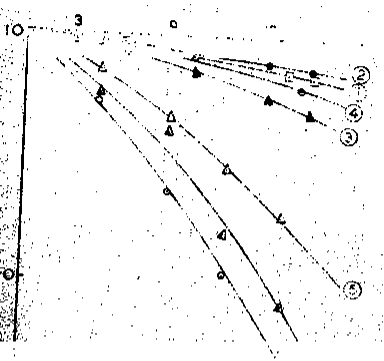


FIG. 5a RESULTS OBTAINED FROM REACTION OF
SMALL DROPLETS WITH
IN AMALGAM AQUEOUS PHASE
PLOTTED
INDICATE
NUMBER IS

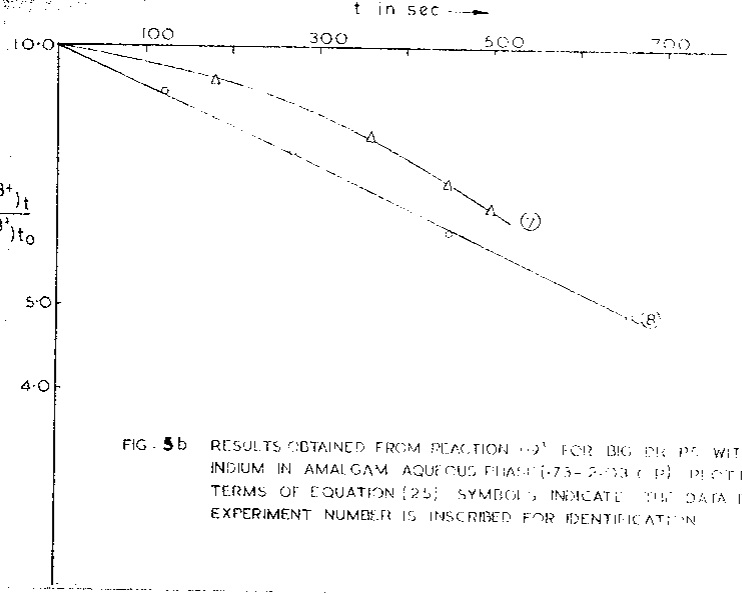


FIG. 5b RESULTS OBTAINED FROM REACTION (2)[1] FOR BIG DR PS WITH [illegible] INDIUM IN AMALGAM. AQUEOUS PHASE (·73–2·03 C P) PLOTTED IN TERMS OF EQUATION (25). SYMBOLS INDICATE TEST DATA POINTS. EXPERIMENT NUMBER IS INSCRIBED FOR IDENTIFICATION.

Polarographic results for the Diffusion co-efficients of ferric ion in the different aqueous media are in table No. III a,b,c.

TABLE III a

In water at 35°C ± .5°C

S.L. No.	Conc. of Fe^{3+} in m.mole per liter.	Diffusion co-efficient in Cm^2/sec.
1	6.25	7.0387×10^{-6}
2	12.50	6.9433×10^{-6}
3	18.75	6.8350×10^{-6}
4	20.00	6.1476×10^{-6}
5	31.25	4.2995×10^{-6}

TABLE III b

In 35 wt% glycerol + water at 35°c ± .5

1	12.5	
2	18.75	2.748×10^{-6}
3	20.00	2.6171×10^{-6}
4	31.25	1.6903×10^{-6}

TABLE III c

In 70 wt percent glycerol at 35°C + .5

1	12.5	6.6145×10^{-7}
2	18.75	5.4300×10^{-7}
3	20.00	7.5238×10^{-7}
4	31.25	4.2540×10^{-7}
5	64.00	1.7012×10^{-7}

Values of Diffusion co-efficient vs concentrations of ferric ion from table III a,b,c had been plotted and shown in Fig.46

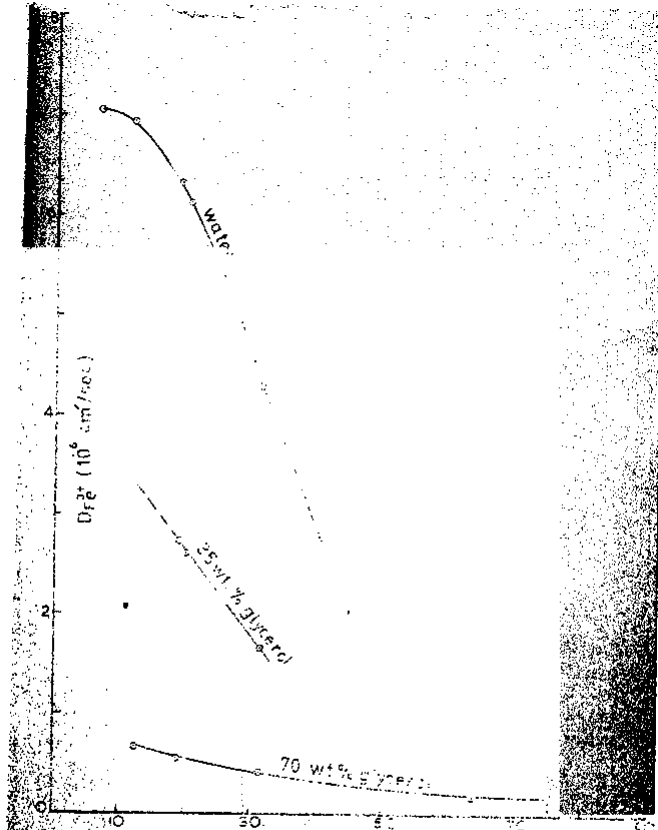


FIG. 6. DIFFUSIVITY OF FERRIC ION AT 35° ± 0.5°C IN VARIOUS GLYCEROL-WATER AND FERRIC ION CONCENTRATIONS.

6. Discussions of the results

A. Kinetic steps.

The reaction being studied is

$$[In]_{amal} + 3(Fe^{3+})_{aq} = (In^{3+})_{aq} + 3(Fe^{2+})_{aq} \qquad (19)$$

the kinetic steps are

1) transport of Fe^{3+} ions to the interface from the aqueous phase (bulk)
2) transport of In from the bulk amalgam to the interface
3) reaction at the interface (eqn.19).
4) transport of Fe^{2+} from the interface to the aqueous phase
5) transport of In^{3+} from the interface to the aqueous phase.

The chemical reaction at the interface (step 3) is expected to be very fast because it essentially involves electron exchange amongst atoms and simple ions. Subramanian and Richardson[25] also found it to be very fast. Hence step(3) is not at all expected to be rate controlling.

If the process is solely transport controlled, then chemical equilibrium amongst the various species may be assumed. The equilibrium relationship for the reaction in equation (19) is as follows:

$$K = \frac{(In^{3+})_i (Fe^{2+})_i^3}{[In]_i \ (Fe^{3+})_i^3} = 3 \times 10^{56} \qquad (22)$$

Where K is the equilibrium constant for the reaction (19). The brackets indicate the concentration of the species in the aqueous and metal phase and the suffix i indicate the concentration of the species at the interface in gram mole per c.c.

Because In^{3+} and Fe^{2+} are moving out from the interface to aqueous phase, $(In^{3+})_i$ and $(Fe^{2+})_i$ must be greater than $(In^{3+})_{bulk}$ and $(Fe^{2+})_{bulk}$ respectively and hence must be greater than zero as soon as the experiment has commenced. Since K is very large, and the numerator in eqn. (22) cannot be infinity, therefore the denominator must be tending to zero. This is possible provided either $[In]_i$ or $(Fe^{3+})_i$ or both tend to zero.

If the rate of transport of Fe^{3+} is less than that of In in the amalgam, then $(Fe^{3+})_i$ would tend to zero and $[In]_i$ would be close to $[In]_{bulk}$. For rate calculations, we may take $(Fe^{3+})_i \simeq 0$ without any significant error. The question is under such circumstances which transport step is rate-controlling ? Step (3) has already been assumed too fast. As regards, transport of Fe^{2+} and In^{3+} [Steps (4) and (5)] are concerned, it may be stated that they are able to adjust themselves to match the rate of transport of Fe^{3+}. The adjustment would be achieved because the values of $(Fe^{2+})_i$ and $(In^{3+})_i$ can go up to their respective limits of solubility. Even if the solubility limits are reached, removal of Fe^{2+} and In^{3+} from the interface is possible by precipitation from solution. In this investigation, no precipitation effect was found. Hence the values of $(Fe^{2+})_i$ and $(In^{3+})_i$ were below their respective solubility limits. Hence

it can be concluded that step (4) and (5) would adjust themselves to match the rate of step (2) and hence step (1) would be the rate-controlling step with $(Fe^{3+})_i \simeq 0$. This has been called "aqueous control" by Subramanian and Richardson[25].

A reverse situation can also be conceived where the rate of transport of In in the amalgam is much slower compared to that of Fe^{3+}. Following the arguments given above, it can be similarly shown that transport of In in the amalgam would be rate-controlling with $[In]_i \simeq 0$. This has been termed as "metal control" by Subramanian and Richardson[25].

A third possibility is mixed-control of step (1) and (2) where the values of both $[In]_i$ and $(Fe^{3+})_i$ tend to zero.

In this investigation, the aim has been to study the aqueous control only and inorder to achieve this the concentration of indium in the amalgam was kept high.

B. Evaluation of Masstransfer co-efficient from experimental data.

Masstransfer coefficients have been calculated from the change of concentration of indium in the aqueous phase. This section presents the various equations that correlated experimental variables with masstransfer coefficients.

Case I : Transport of Fe^{3+} in aqueous solution rate-controlling

If the number of droplets per unit time be n per sec. and the residence time of one droplet in aqueous column be τ sec, then the number of droplets in the bath at any moment would be equal to $n\tau$.

Since the transport of Fe^{3+} is rate controlling and $(Fe^{3+})_i \simeq 0$, hence

$$J_{Fe^{3+}} = K_{Fe^{3+}} C_{Fe^{3+}} \qquad (23)$$

Hence by molal balance of (Fe^{3+}) in the aqueous solution

$$n\pi d^2 K_{Fe^{3+}} C_{Fe^{3+}} dt = - Va\, dc_{Fe^{3+}} \qquad (24)$$

Where $J_{Fe^{3+}}$ = mass flux of ferric ion

d = diameter of the droplets in cm.

$K_{Fe^{3+}}$ = mass transfer co-efficient for ferric ion in cm. sec^{-1}.

$C_{Fe^{3+}}$ = concentration of ferric ion in the aqueous phase whose transfer was rate controlling in m. moles/liter

t = time in sec.

Va = volume of the aqueous phase in c.c.

By rearrangements and integrating within the limits it could be expressed as

$$\log \frac{(Fe^{3+})_t}{(Fe^{3+})_{t_o}} = - K't. \qquad (25)$$

Where subscript t_o and t denote the concentration on initial and final times and $K' = \dfrac{n\pi d^2 K_{Fe^{3+}}}{2.303\ Va}$

From the plots in figure 5a, 5b, K' being the slope and $K_{Fe^{3+}}$ was calculated from the experimental values of n,d,va etc as mentioned in the experimental chapter. The concentration of In^{3+} in the aqueous phase at different interval of times were determined polarographically and the concentrations of ferric ion were calculated from stoichiometry as in eqn (19) (for details refer to chapter 4.)

Logarithm of concentration ratio of ferric ion can be plotted against the weight of amalgam passed per unit time instead of the case as shown in Fig. 5a, 5b. The total weight of amalgam passed during the course of the experiment over the period of t sec will be

$$n \frac{\pi}{6} d^3 t \times 13.6 = W \qquad (26)$$

Where W is the weight of amalgam in gms.

From eqn. (25)

$$- K't = - \frac{n \tau \pi d^2 K_{Fe^{3+}}}{2.303 V_a} t \qquad (27)$$

Substituting eqn. (26) in (27) and rearrangements

$$-K't = - \frac{.194 \tau K_{Fe^{3+}}}{d V_a} W \qquad (28)$$

Hence eqn. (25) is modified and expressed as

$$\log \frac{(Fe^{3+})_{final}}{(Fe^{3+})_{initial}} = - .194 \tau \frac{K_{Fe^{3+}}}{dV_a} W \qquad (29)$$

Plot $\log \frac{(Fe^{3+})_{final}}{(Fe^{3+})_{initial}}$ VS W should be straight line and from the slope $K_{Fe^{3+}}$ could be calculated when the other terms like d, τ, V_a etc. were experimentally determined.

Case 2: Transport of In in the amalgam rate controlling.

The masstransfer co-efficient for indium K [In] could be calculated similary by mass balance as before. The concentration of indium in the drop phase would be constant in each experiment. Since the transport of indium is rate controlling and $[In]_i \approx 0$ hence

$$J_{In} = K_{[In]} C_{[In]} \qquad (30)$$

The indium balance equation expressed as

$$\pi d^2 K_{[In]} C_{[In]} dt = - V_m dc_{[In]} \qquad (31)$$

Where $K_{[In]}$ = mass transfer co-efficient for indium in drop phase.

$C_{[In]}$ = conc. of In in the amalgam forming the droplets in weight percent.

V_m = volume of the drop in c.c.

t = time in sec.

d = diameter of the drop in cm.

By rearrangements and integrating within the limits

$$\log \frac{[In]_t}{[In]_{t_o}} = -K''t \qquad (32)$$

Where the subscripts t_o, t denotes the concentrations in initial and final time and $K'' = \frac{\pi d^2 K_{[In]}}{2.303\ V_m}$

From the plot $\log \frac{[In]_t}{[In]_{t_o}}$ VS t, the slope is K", hence $K_{[In]}$ can be calculated.

Concentration of indium after an interval of time can be obtained either by analysing the used amalgam or by analysing its concentration in the aqueous phase by polarography.

C. Nature of masstransfer data

The results with reaction in eqn. (19) with 0.05% indium in amalgam and ferric ion varying between 10-75m. moles per liter in the aqueous phase (.73-11.75 cp) are plotted in fig - 5a, 5b where $\log \frac{(Fe^{3+})_t}{(Fe^{3+})_{t_o}}$ are shown as function of time.

TABLE -4

The different values calculated from Fig. 5a, 5b by computer

Exp. No.	$(Fe^{3+})_t$ in m.moles per liter on exp.points	$D_{Fe^{3+}}$ in $cm^2/sec.$	$K_{Fe^{3+}}$ in cm/sec.	$K_{Fe^{3+}} / D^{\frac{1}{2}}_{Fe^{3+}}$	$K_{Fe^{3+}}$ average / $D^{\frac{1}{2}}_{Fe^{3+}}$ average
	24.6425	4.75×10^{-7}	1.193×10^{-2}	17.30	
1	24.245	4.50×10^{-7}	1.243×10^{-2}	21.30	
	23.8075	4.84×10^{-7}	1.493×10^{-2}	21.50	22.377
	22.610	4.96×10^{-7}	1.853×10^{-2}	26.50	
	22.335	5.00×10^{-7}	1.928×10^{-2}	27.40	
2	-	1.25×10^{-7}	1.04×10^{-2}	29.4	29.4
	9.530	7.45×10^{-7}	2.13×10^{-2}	24.7	
3	9.425	7.60×10^{-7}	3.03×10^{-2}	34.9	38.61
	8.436	7.85×10^{-7}	3.94×10^{-2}	44.5	
	8.002	7.97×10^{-7}	4.47×10^{-2}	50.1	
	24.475	4.77×10^{-7}	1.202×10^{-2}	17.55	
4	23.63	4.86×10^{-7}	1.722×10^{-2}	24.80	
	22.6425	4.96×10^{-7}	2.24×10^{-2}	31.60	27.86
	21.6625	5.12×10^{-7}	2.66×10^{-2}	37.30	
	20.5775	6.05×10^{-6}	1.245×10^{-1}	50.60	
	16.20	6.68×10^{-6}	1.502×10^{-1}	58.20	
5	12.955	6.93×10^{-6}	1.706×10^{-1}	64.80	61.60
	10.07	7.02×10^{-6}	1.911×10^{-1}	72.40	
	22.5625	2.42×10^{-6}	6.4×10^{-2}	41.30	
	19.7275	2.65×10^{-6}	8.12×10^{-2}	49.90	
6	17.3175	2.85×10^{-6}	9.49×10^{-2}	56.50	52.37
	14.8975	3.12×10^{-6}	10.8×10^{-2}	61.20	
	22.925	2.4×10^{-6}	6.27×10^{-2}	40.45	
	19.43	2.68×10^{-6}	10.19×10^{-2}	62.24	
7	17.3625	2.85×10^{-6}	12.15×10^{-2}	71.58	63.19
	16.2975	2.94×10^{-6}	13.13×10^{-2}	77.10	
8	-	6.46×10^{-6}	1.325×10^{-1}	52.18	52.18
	20.9725	6.00×10^{-6}	1.25×10^{-1}	51.019	
9	x	x	x	x	59.12
	14.2425	6.85×10^{-6}	1.584×10^{-1}	60.53	
	11.6475	6.98×10^{-6}	1.73×10^{-1}	65.481	

The derived masstransfer co-efficients are given in table- 4.

If the aqueous control prevails, then a change in the concentration of indium in amalgam should not affect the rate of masstransfer. In order to verify this, the concentration of indium was increased by a factor of 3 in two experiments. In one case the rate increased slightly (Compare $K_{Fe}3+$ values for expt. No. 3 and 9) and can be ignored. In the other case (Compare $K_{Fe}3+$ values of expt. No.1 and 4), the rate increased by approximately 20 percent. However it may be partly due to some decrease in drop size in expt. No.4. Therefore, it may be safely concluded that the influence on rate because of a change in C_{In} is not significant. The little variation may be due to some other surface effect because of higher concentration of indium in the amalgam.

Further more, according to eqn.(25), $\log \frac{(Fe^{3+})_t}{(Fe^{3+})_{t_o}}$ should vary linearly with t and should be independent of initial (Fe^{3+}) concentration provided other factors such as drop size, viscosity etc. are kept constant. At 70 wt percent glycerol and small drop size, several experiments were conducted at ferric ion concentrations of 10 , 25 and 75 m. moles. The table No.4 shows the mass transfer co-efficient decreases systematically with increase in concentrations. However, it has to be recognised that part of this variation is a result of decrease of diffusion co-efficient with increase in concentration. Fig.6 shows the variation of diffusion co-efficient of ferric ion with change in concentration in different aqueous medium.

If all other factors remain constant, then $K_{Fe}3+/ D^{\frac{1}{2}}_{Fe}3+$ should be constant[4]. Table 4 compare $K_{Fe}3+/D^{\frac{1}{2}}_{Fe}3+$ and it can be seen that $K_{Fe}3+/D^{\frac{1}{2}}_{Fe}3+$ does neither vary much nor vary systematically for expt. 1-4. Therefore, from this point of view, the experiments 1-4 do not basically violate the aqueous control condition.

Figures 5a, 5b show that the curves mostly show some systematic deviation from linearity. Computer programming with various orders indicated that second order polynomial gives best fitting with the experimental results. The slopes and the masstransfer coefficients at various intervals were calculated by computer and given in table 4. The table shows that $K_{Fe}3+$ increases with time except for experiment No.2 and 8 where $K_{Fe}3+$ remains constant. Again this deviation could be due to variation of $D_{Fe}3+$ with concentration. Therefore, $K_{Fe}3+/D^{\frac{1}{2}}_{Fe}3+$ should be the real basis for deciding the rate controlling step. As table 4 indicates, variation of $D_{Fe}3+$ alone cannot account for the variation in $K_{Fe}3+$ completely. It is not possible to offer any concrete explanation at this stage, but some concentration dependent factors seems to be playing a role or there may be surface effects.

While comparing the results for experiments 5 and 6, it can be seen that masstransfer coefficient values in exp-5 has increased by a factor of 2 than in exp. No.6, though lower terminal velocity and higher residence time has been observed in the later case. The results in table 4 indicates the variation of diffusion coefficient due to change in aqueous medium

must have contributed in the enhancement of the masstransfer coefficient values. In the case of experiment 7 and 8 with bigger drop size and in the same medium as in exp. 5 and 6, the mass transfer coefficient values has differed only by 21 percent, even though the diffusion coefficient values are very much different. Photographic studies as indicated earlier revealed that the bigger droplets oscillated vigorously during its passage through the aqueous phase and no droplet oscillation was there in the case of smaller one. Since the residence time and terminal velocity in exp. 7 and 8 had not differed much, it may be concluded that the droplet oscillation had influenced the masstransfer coefficient values.

D. Correlation of experimental data.

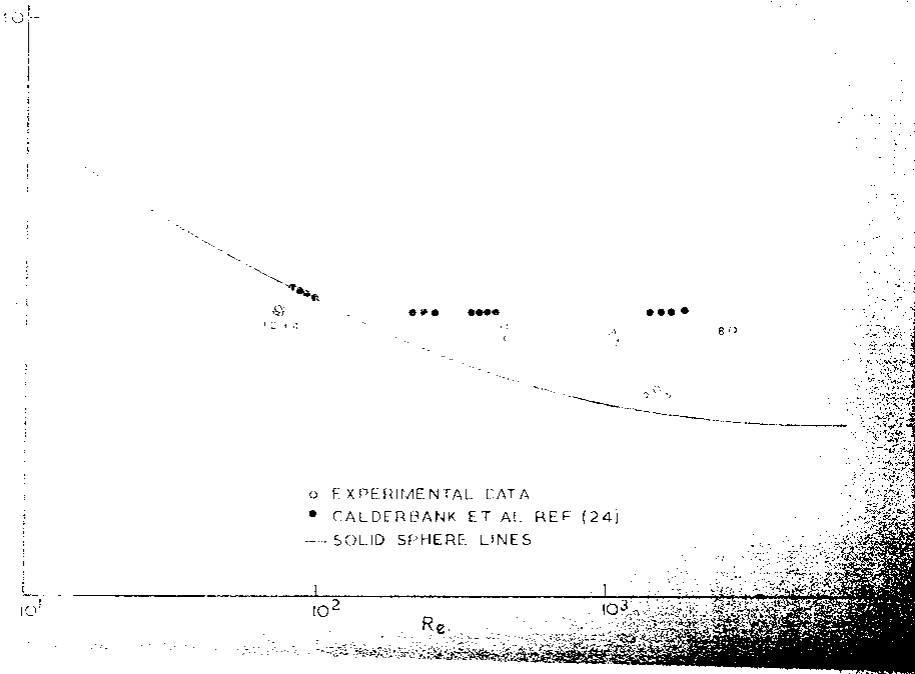
Reynolds number, drag coefficient, Schmidt number, Sherwood number have been calculated from the data in tables 1a, 1b and using average diffusion and masstransfer coefficient from experimental results. The values are given in table 5.

Experimental drag coefficient values have been plotted against Reynolds number in log scale in fig.7 and compared with solid sphere behaviours. In the case of small drops in high viscosity fluid and in the range of Re below 100 as in the case of experiments 1-4, the deviation from solid sphere line is not much. This is in conformity with the photographic observations that small drops did not oscillate and thus behaved like solid spheres. In the case of experiment 5 and 9 with small droplets in water, the slight deviation from solid sphere line is there which indicated slight oscillation in the droplet. The point 6,7,8 matches with results of Calderbank[24] et al

TABLE - 5

Calculated from table No.1a, 1b and plots 5a, 5b

Exp.	Average diffusion coeff. of Fe^{3+} in $cm^2/sec.$	Average masstransfer coeff. in cm/sec.	N_{Re}	C_D	N_{Sc}	N_{Sh}
1	4.87×10^{-7}	1.562×10^{-2}	75.5	.975	2.07×10^{5}	5.12×10^{3}
2	1.25×10^{-7}	1.04×10^{-2}	76.0	.955	7.95×10^{5}	1.33×10^{4}
3	7.72×10^{-7}	3.392×10^{-2}	75.0	.970	1.30×10^{5}	6.95×10^{3}
4	4.93×10^{-7}	1.956×10^{-2}	73.5	.945	2.025×10^{5}	6.15×10^{2}
5	6.67×10^{-7}	1.691×10^{-1}	1520	.53	1.09×10^{3}	3.77×10^{3}
6	2.76×10^{-6}	8.7×10^{-2}	445	.865	6.75×10^{3}	5.03×10^{3}
7	2.72×10^{-6}	10.42×10^{-2}	1080	.815	6.85×10^{3}	10.85×10^{3}
8	6.46×10^{6}	1.325×10^{-1}	2820	.852	1.12×10^{3}	5.79×10^{3}
9	6.61×10^{-6}	1.52×10^{-1}	1520	.53	1.10×10^{3}	3.64×10^{3}


o EXPERIMENTAL DATA
• CALDERBANK ET AL REF (24)
— SOLID SPHERE LINES
10^2
10^3
R_e

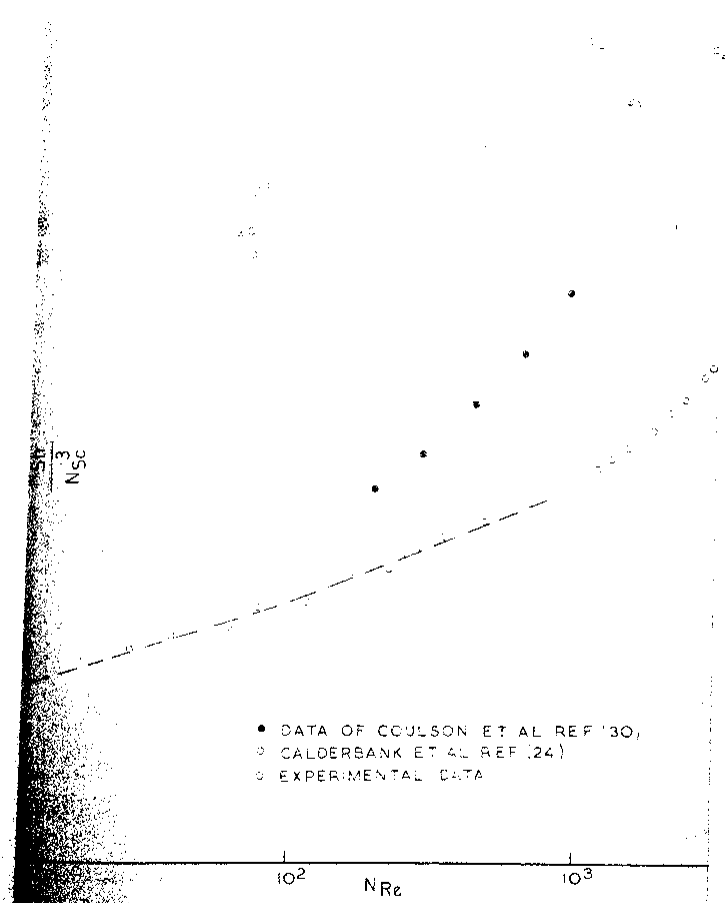


FIG. 8 CONTINUOUS PHASE MASSTRANSFER COEFFICIENT CORRELATION FROM AMALGAM DROP DATA

within the limits of experimental errors. Photographic records revealed substantial oscillation in these cases.

Heat transfer from droplets to a continuous phase with two liquid system had been studied by Calderbank[24] and Coworkers. The relationship obtained by them is presented in fig.8 along with those of Coulson[30] et al. Calderbank[24] et al had a situation similar to present investigation in the sense that they employed falling mercury droplets in aqueous-glycerol mixture and studied the aqueous control case. Hence in the similar manner the experimental data of the present investigation have also been plotted in fig.8.

However, as figure 8 show, the data from this investigation are approximately an order of magnitude higher than those of Calderbank et al[24]. The results of Coulson[30] et al deviate from the other and appears to merge with the present data when the former is extrapolated to higher Re values.

A comparisdn with heat transfer from solid sphere reveals that at low Reynolds no, circulation inside the droplet does not have any appreciable effect on continuous phase mass-transfer coefficient. According to Calderbank et al, their data deviated from the solid sphere behaviour above Re= 200 presumably due to droplet oscillation. If this interpretation is correct then it is to be noted that the droplet oscillation or any other causes depending upon the system will result in the departure from solid sphere behaviour and may have entirely different effect on the masstransfer coefficient. The higher positive departure of Coulson's results cannot be otherwise explained. Following the same argument it may be said that perhaps the droplet oscillation began at much lower Re in the present

investigation because of large lowering of surface tension as a result of high masstransfer rate. Also a high masstransfer rate may by itself change the correlation.

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7. Suitable Model System

It is strongly felt that in Blast furnace the desulfurization of molten metal over and above that at interface takes place to a large extent when metal droplets passed through the thick layer of slag. In order to study the possibities of mass-transfer from the droplets a suitable system had been selected and the model was studied under room temperature conditions. All the experiments were conducted at 35° ± .1°C by passing mercury-indium amalgam droplets of different sizes through aqueous media having various concentration of glycerol and ferric ion. Indium concentrations in aqueous phase were determined by polarographic technique, then by suitable stoichiometry the depletion of ferric ion in the aqueous media were obtained. By suitable relationship masstransfer coefficients were calculated. Diffusion coefficient of ferric ion in the different aqueous media were measured. Terminal velocity and residence time for the droplets in the aqueous phase were determined by photographic technique.

The conclusions of the present investigations are summarized as follows:

1) Appreciable amount of masstransfer could not be measured even by polarography when the experiments were conducted at 21-22°C in the present apparatus. By increasing the column height, the residence time for the droplets in the aqueous phase could be increased and appreciable amount of masstransfer could be measured at 21-22°C. With a short column, experiments at higher temperature solved the problem.

2) Systematic studies revealed that aqueous control condition existed in all experiments.

3) Experiments with 70 wt percent glycerol in aqueous phase revealed that the value of masstransfer coefficient decreases progressively with increase in the concentration of ferric ion. This could be attributed to lowering of diffusion coefficient due to increase of ferric ion concentration.

4) In the case of small droplets, the masstransfer coefficient decreased with increase in the viscosity of the aqueous phase whereas with bigger droplets, the difference in the masstransfer coefficient values were small because of severe droplet oscillation during their passage through the aqueous phase.

5) Smaller droplets behaved more or less as solid sphere while passing through high viscosity medium whereas in low viscosity media, oscillation to some extent could be observed. In the case of bigger droplets, severe oscillation in the different viscous media had been observed.

6) Terminal velocity and residence time for all the droplets in the different aqueous media had been determined and systematic variations had been observed.

7) Photographic studies revealed that the smaller droplets were reaching their terminal velocity within 1 cm from the top of liquid surface where as 3-3.5 cm in the case of bigger one.

8) The masstransfer coefficients were one order magnitude higher than that expected from solid sphere behaviour presumably because of droplet oscillations and possibly the Marangani effect.

NOMENCLATURE

A	=	area in cm^2
Ae	=	area of a sphere whose volume is equal to the volume of ellipsoid in cm^2.
C	=	concentration in gm mole/cc.
CC	=	concentration in the bulk
D	=	diffusivity in cm^2/sec.
d_1	=	diameter of the drop in cm
de	=	equivalent diameter in cm
d_H	=	diameter in horrizontal direction in cm
d_V	=	diameter in vertical direction in cm.
E	=	ecentricity of drop $\frac{d_H}{d_V}$
E'	=	ecentricity of drop $\frac{d_1}{d_e}$
g	=	gravitational acceleration cm/sec^2
K_D	=	over all masstrasfer co-efficient cm/sec.
k_D	=	drop phase masstransfer coefficient cm/sec.
k_C	=	continuous phase masstransfer coefficient
U	=	terminal velocity cm/sec.
U_L	=	terminal velocity for liquid drop
U_S	=	terminal velocity for solid sphere from Stokes law.
r	=	radius of the droplet in cm
N_{Re}	=	Reynolds number $\frac{\rho U d}{\mu}$
N_{Sc}	=	Schmidt number $= \frac{\mu}{\rho D}$

N_{Sh} = Sherwood number $\frac{Kd}{\mathcal{D}}$

n = parameter

Greek letters

π = 3.1416

ρ_c = density of continuous phase

ρ_D = density for drop phase gm/cc

σ = surface tension dynes/cm

ω = frequency of oscillation per sec.

μ_c, μ_D = continuous and drop phase viscosity in poise.

δ = diffusion boundary layer thickness in cm.

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